

Effect of Chemical Structure on the Electrochemical Cleavage of Alkoxyamines

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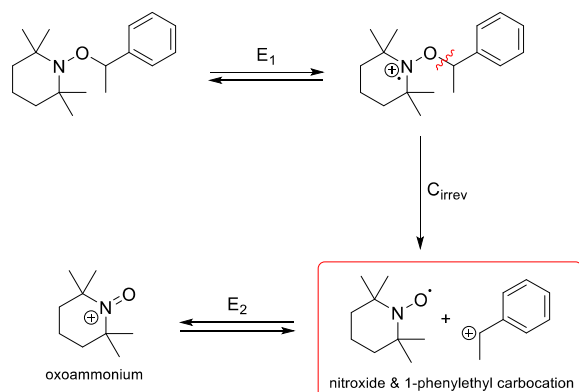
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ABSTRACT: A test set of 14 TEMPO-based alkoxyamines was studied *via* a combination of cyclic voltammetry (CV) and accurate quantum chemistry to assess the effect of substituents on electrochemical cleavage. The experimental oxidation potentials of the alkoxyamines fell into the range of 1.1–1.6 V versus Ag/AgCl in acetonitrile, were well reproduced by theory (MAD 0.04V), with values showing good correlation with the σ_R Hammett parameters of both the R group and the OR group in TEMPO-R. Importantly, most of the studied alkoxyamines underwent oxidative cleavage to form either TEMPO• and R• or TEMPO⁺ and R•, with the former favored by electron donating substituents on R (e.g., 2-oxolane, Ac, CH(CH₃)Ph, *i*-Pr, *t*-Bu) and the latter by electron withdrawing substituents (Bn, allyl, CH(CH₃)C(O)OCH₃, C(CH₃)₂C(O)OCH₃, CH(CH₃)CN). Where R is not stabilized (e.g. R = CH₂C(O)OCH₃, Me, Et), fully or almost fully reversible oxidation – without cleavage – was observed, making these species promising candidates for battery applications. Finally, in the case of R = Ph where N–O cleavage occurred, a phenoxy cation and an aminyl radical were generated. Based on these results, TEMPO-based alkoxyamines can provide a variety of electrochemically generated carbon-centered radicals and carbocations for use in synthesis, polymerization and surface modification.

INTRODUCTION

Carbon-centered cations and radicals are common reactive intermediates in organic reactions, however their application in complex molecular synthesis is frequently limited by their preparative methods.¹ Conventional approaches often require strong acids,² energetic conditions,³ or are applicable for the generation of only the most stabilized ions.^{4–5} The use of electrochemical methods in organic synthesis, which is currently undergoing a renaissance,⁶ may offer a potential solution to this problem.⁷



Scheme 1. Electrochemical oxidation and subsequent cleavage of 2,2,6,6-tetramethyl-1-(1-phenylethoxy)piperidine (or TEMPO-CH(CH₃)Ph) *via* an EC_{irrev}E mechanism.

Recently, we showed that upon one electron oxidation, 2,2,6,6-tetramethyl-1-(1-phenylethoxy)piperidine (denoted here as TEMPO-CH(CH₃)Ph) underwent irreversible mesolytic cleavage to the corresponding nitroxide radical (TEMPO) and the 1-phenylethyl carbocation (Scheme 1). The TEMPO radical then underwent a second (reversible) one electron oxidation, resulting in an overall EC_{irrev}E mechanism. The results were confirmed using a combination of cyclic voltammetry (Figure 1), electron paramagnetic resonance (to confirm nitroxide generation) and computational chemistry (to further confirm the preferred cleavage pathway).⁸

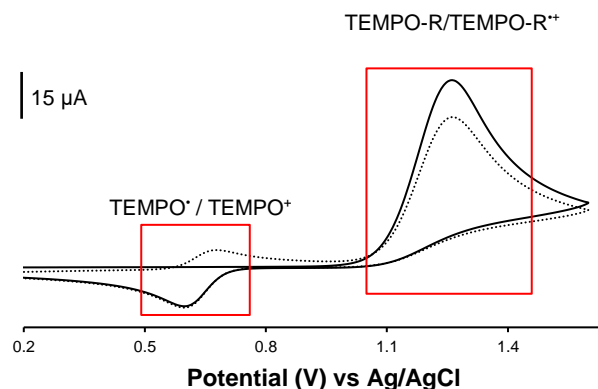
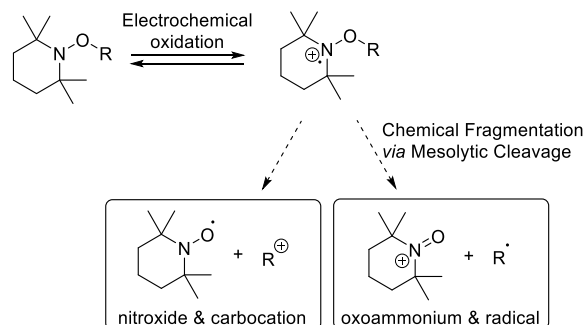


Figure 1. Cyclic voltammogram of TEMPO-CH(CH₃)Ph (2 mM) in Bu₄NClO₄/MeCN (0.1 M), sweep rate 100 mV s⁻¹. Solid line: scan 1. Dotted line: scan 2.

On the first forward scan of the cyclic voltammogram in Figure 1, the TEMPO oxidation peak is absent but there is a large irreversible peak for the oxidation of TEMPO-CH(CH₃)Ph, which undergoes fragmentation to the ⁺CH(CH₃)Ph carbocation and the TEMPO radical. The peak also has a contribution from the immediate oxidation of the resulting TEMPO^{•+} fragment to TEMPO⁺. On the reverse scan, reduction of TEMPO⁺ to TEMPO[•] is visible and the TEMPO-CH(CH₃)Ph oxidation peak is smaller reflecting the consumption of the alkoxyamine. Subsequent work showed that when tethered to a Si(100) surface, the carbocation can trap ferrocenemethanol, thus providing proof-of-concept for use of this technology in surface modification.⁹



Scheme 2. Electrochemical oxidation and subsequent cleavage to either a nitroxide and carbon-centered cation, or an oxoammonium and carbon-centered radical.

While mesolytic cleavage to a nitroxide radical and carbocation was confirmed for TEMPO-CH(CH₃)Ph, in principle two cleavage pathways are possible: (1) formation of a nitroxide and carbon-centered cation, or (2) an oxoammonium cation and carbon-centered radical (Scheme 2). Alteration of the nitroxide or R-group substituent should favor one fragmentation pathway over another, or could inhibit fragmentation altogether. In the latter case, a simple reversible oxidation would be observed with potential applications in battery technologies.¹⁰ When fragmentation occurs, the *in situ* and “on-demand” production of carbocations (or radicals) *via* electrochemically induced fragmentation gives rise to a plethora of potential applications and expands upon traditional methodologies in the organic synthesis of small molecules, and tandem polymerizations.⁷

In the present work we use a combination of cyclic voltammetry and accurate computational chemistry to study the effect of the R-group on the electrochemical oxidation and cleavage of TEMPO-R. Our aim is to explore the scope of carbocations that can be generated in this way and establish under what conditions this approach can be used to generate carbon-centered radicals instead.

EXPERIMENTAL

Computational Procedures. All geometry optimizations were performed using the M06-2X functional and 6-31+G(d,p) basis set.^{11–15} All geometries were confirmed to be local minima (possessing no imaginary frequencies). High-level single-point energies were calculated using the composite *ab initio* G3(MP2,CC)(+)//M06-2X/6-31+G(d,p) method, a variation of standard G3(MP2,CC)¹⁶ where MP2 and CCSD(T) single point calculations with the 6-31G(d) basis set are replaced with 6-31+G(d) and MP2/6-31G(d) geometries and frequencies are

substituted with those calculated at M06-2X/6-31+G(d,p). These high-level calculations were utilized in conjunction with the ONIOM style approximation for larger systems,^{17–18} where the full system was modelled using G3(MP2,CC)//M06-2X/6-31+G(d,p). Solvation corrections were performed using a thermocycle approach, using COSMO-RS.^{19–21} A full systematic conformational search was performed on species investigated in this work, with a resolution of 120° around sp³-sp³ and sp³-sp² bonds and 180° around sp²-sp² bonds. All standard *ab initio* molecular orbital theory and density functional theory (DFT) calculations were carried out using Gaussian 09,²² except for CCSD(T) calculations which were performed with the Molpro 2015²³ software package. The ADF 2014 package²⁴ was used to compute COSMO-RS solvation free energies at the BP/TZP level of theory (as it was parameterized for), and the remaining parameters were kept as default values. Oxidation potentials (*E*_{ox}) were computed using eqn 1:

$$E_{ox} = \frac{-\Delta G_{sol}}{nF} \quad (1)$$

Where, *n* is number of electrons transferred (*n* = 1 in this case) and *F* is the Faraday constant. This *E*_{ox} was then referenced against the value of the Ag/AgCl couple in acetonitrile (see the Supporting Information for more details). For more details regarding the theoretical prediction of oxidation potentials, the reader is directed to Ho *et al.*²⁵

Cyclic Voltammetry. Electrochemical experiments were carried out using a PalmSens4 Potentiostat/Galvanostat/Impedance Analyzer and a single compartment three-electrode glass cell. Electrochemical experiments were performed at room temperature (23 ± 2 °C) under argon. Glassy carbon disks served as the working electrode, a platinum coil was used as counter electrode, and silver/silver chloride as the reference electrode. The active area of the working electrode (0.060–0.065 cm²) was electrochemically-determined from refinement of an E model against experimental voltammograms measured in 7.1 × 10^{−2} M MeCN/Bu₄NClO₄ in the presence of 1.0 × 10^{−3} M ferrocene. Alkoxyamine (ca. 2 mM, analytical concentrations are listed in Fig. S1 to S14) was dissolved in 1.0 × 10^{−1} M MeCN/Bu₄NClO₄. The solution was stirred and bubbled briefly under argon prior to measurement. All cyclic voltammograms were obtained after bubbling, with no stirring. Digital simulations of cyclic voltammetry were performed in DigiElch-Professional v7 (ElchSoft). Butler-Volmer kinetics was used to estimate charge transfer parameters and we assumed a semi-infinite 1D diffusion. The cell *iR* drop was not compensated during measurement. Values of cell resistance were measured by electrochemical impedance spectroscopy and used in the simulations (data of resistance were between 140 and 160 Ω for MeCN/Bu₄NClO₄ at a concentration of 1.0 × 10^{−1} M). Adsorption steps were neglected in the simulations and the transfer coefficients for both electron transfer reactions (TEMPO/oxoammonium and alkoxyamine/alkoxyamine^{•+}) were assumed as constants (α, 0.5) for fitting purposes. In the simulation the homogeneous chemical step for the fate of the putative benzylic cation fragment (R⁺) was not considered as this has no effect on the quality of the fits. For the chemical step, only the forward (*k*_f) constant is considered. The second-order backward (*k*_b) constant tends to zero (i.e. backward chemical reaction is not operative in the time scale of the experiments).

General Synthetic Procedures. All reactions were completed under anaerobic conditions unless otherwise stated. Under these conditions, standard Schlenk and vacuum line techniques were used with the addition of dried and degassed solvents. Brine refers to a saturated aqueous solution of NaCl. Petrol refers to the fraction of petroleum ether boiling in the range 40–60 °C. Starting materials were purchased from commercial sources unless otherwise stated. Flash column chromatography was carried out using Merck silica gel 60 (0.04–0.063 mm, 230–400 mesh). All NMR spectra were obtained at 298 K on a Bruker AVANCE 400 (¹H at 400.1 MHz and ¹³C at 100.5 MHz). Chemical shifts (δ) are reported in ppm and referenced in accordance to CHCl₃ (δ_H 7.26) or CDCl₃ (δ_C 77.16). HRMS data was obtained on an Orbitrap QE plus or LCT Premier mass spectrometer.

1-Methoxy-2,2,6,6-tetramethylpiperidine A1 A solution of 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO) (1.0 g, 6.4 mmol) and naphthalene (16 mg, 0.12 mmol, 2 mol%) in THF (2 mL) was added to a Schlenk flask containing sodium (200 mg, 8.7 mmol, 1.4 eq.) in THF (2 mL). The reaction mixture was stirred at room temperature until the color changed from orange to green (4 h). The resulting solution was added to a solution of iodomethane (600 μL, 9.6 mmol, 1.5 eq.) in THF (2 mL), then stirred at room temperature for 18 hours. Diluted with saturated NH₄Cl solution (10 mL), extracted with ethyl acetate (3 × 10 mL), dried over Na₂SO₄ and concentrated under reduced pressure. Purified by flash chromatography, eluting with 5% EtOAc–petrol, to give 1-methoxy-2,2,6,6-tetramethylpiperidine **A1** (520 mg, 47%) as a colorless oil; ¹H NMR (400 MHz; CDCl₃) 3.57 (s, 3H), 1.54–1.30 (m, 6H), 1.13 (s, 6H), 1.04 (s, 6H); ¹³C NMR (100 MHz, CDCl₃) 65.4 (CH₃), 59.8 (C), 39.8 (CH₂), 33.1 (CH₃), 20.1 (CH₃), 17.2 (CH₂); HRMS *m/z* calculated for C₁₀H₂₂NO⁺ (M + H)⁺ 172.1696, found 172.1696 (+ 0.227 ppm error).

1-Ethoxy-2,2,6,6-tetramethylpiperidine A2 A solution of 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO) (1.0 g, 6.4 mmol) and naphthalene (16 mg, 0.12 mmol, 2 mol%) in THF (4 mL) was added to a Schlenk flask containing sodium (200 mg, 8.7 mmol, 1.4 eq.) in THF (4 mL). The reaction mixture was stirred at room temperature until the color changed from orange to dark green (1–2 h). The resulting solution was added to a solution of iodoethane (770 μL, 9.6 mmol, 1.5 eq.) in THF (2 mL) at 0 °C, then stirred at room temperature for 18 hours. Diluted with saturated NH₄Cl solution (20 mL), extracted with ethyl acetate (3 × 20 mL), dried over Na₂SO₄ and concentrated under reduced pressure. Purified by flash chromatography, eluting with 5% EtOAc–petrol, to give 1-ethoxy-2,2,6,6-tetramethylpiperidine **A2** (421 mg, 36%) as a colorless oil; ¹H NMR (400 MHz; CDCl₃) 3.72 (q, *J* = 7.2 Hz, 2H), 1.53–1.25 (m, 6H), 1.10 (s, 6H), 1.07 (t, *J* = 7.2 Hz, 3H), 1.06 (s, 6H); ¹³C NMR (100 MHz, CDCl₃) 72.1 (CH₂), 59.4 (C), 39.6 (CH₂), 33.0 (CH₃), 20.0 (CH₃), 17.1 (CH₂), 13.5 (CH₃); HRMS *m/z* calculated for C₁₁H₂₄NO⁺ (M + H)⁺ 186.1852, found 186.1862 (+ 2.10 ppm error).

1-(*iso*-Propoxy)-2,2,6,6-tetramethylpiperidine A3²⁶ A solution of 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO) (797 mg, 5.1 mmol) and CuBr (21 mg, 0.15 mmol, 3 mol%) in *t*-BuOH (5 mL) was added stirred at 60 °C under N₂. To this solution was added isobutyraldehyde (2.0 mL, 22 mmol, 4.3 eq) followed by slow addition of hydrogen peroxide (1.2 mL, 30% v/v, 15 mmol, 3.0 eq). The reaction mixture was stirred at room

temperature until a color change from orange to blue was observed. Diluted and extracted with ethyl acetate (3 × 10 mL), washed with sat. NH₄Cl solution (2 × 10 mL) and brine (2 × 10 mL), dried over Na₂SO₄ and concentrated under reduced pressure. Purified by flash chromatography, eluting with 10% EtOAc–petrol, to give 1-(*iso*-propoxy)-2,2,6,6-tetramethylpiperidine **A3** (595 mg, 59 %) as a colorless oil; ¹H NMR (400 MHz; CDCl₃) 3.98 (sept, *J* = 6.2 Hz, 1H), 1.45–1.10 (m, 6 × CH₃ and 3 × CH₂, 24H); ¹³C NMR (100 MHz; CDCl₃) 75.1 (CH), 59.5 (C), 40.3 (CH₂), 34.5 (CH₃), 22.4 (CH₃), 20.3 (CH₃), 17.4 (CH₂); HRMS *m/z* calculated for C₁₂H₂₆NO⁺ (M + H)⁺ 200.2009, found 200.2010 (+ 0.345 ppm error).

1-(*tert*-Butoxy)-2,2,6,6-tetramethylpiperidine A4²⁶ A solution of 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO) (743mg, 4.76 mmol) and CuBr (20 mg, 0.14 mmol, 3 mol%) in *t*-BuOH/H₂O (6:1, 7 mL) was stirred at room temperature. To this solution was added trimethylacetaldehyde (2.1 mL, 19 mmol, 4.1 eq) followed by slow addition of hydrogen peroxide (1.5 mL, 30% v/v, 4.8 mmol, 4.0 eq). The reaction mixture was stirred at room temperature until a color change from orange to blue was observed. Diluted and extracted with ethyl acetate (3 × 20 mL), washed with sat. NH₄Cl solution (3 × 10 mL) and brine (3 × 10 mL), dried over MgSO₄ and concentrated under reduced pressure. Purified by flash chromatography, eluting with 10% EtOAc–petrol, to give 1-(*tert*-butoxy)-2,2,6,6-tetramethylpiperidine **A4** (620 mg, 61 %) as a colorless oil; ¹H NMR (400 MHz; CDCl₃) 1.52–1.42 (m, 6H), 1.28 (s, 9H), 1.12 (s, 6H), 1.07 (s, 6H); ¹³C NMR (100 MHz; CDCl₃) 59.2 (C), 41.0 (CH₂), 35.0 (CH₃), 29.6 (CH₃), 20.6 (CH₃), 17.3 (CH₂); HRMS *m/z* calculated for C₁₃H₂₈NO⁺ (M + H)⁺ 214.2165, found 214.2171 (+ 0.00 ppm error).

1-(Allyloxy)-2,2,6,6-tetramethylpiperidine A5 A solution of 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO) (1.0 g, 6.4 mmol) and naphthalene (32 mg, 0.25 mmol, 4 mol%) in THF (4 mL) was added to a Schlenk flask containing sodium (200 mg, 8.7 mmol, 1.4 eq.) in THF (4 mL). The reaction mixture was stirred at room temperature until the color changed from orange to dark green (1–2 h). The resulting solution was added to a solution of allyl bromide (830 μL, 9.6 mmol, 1.5 eq.) in THF (2 mL) at 0 °C, then stirred at room temperature for 18 hours. Diluted with saturated NH₄Cl solution (20 mL), extracted with ethyl acetate (3 × 20 mL), dried over Na₂SO₄ and concentrated under reduced pressure. Purified by flash chromatography, eluting with 10% EtOAc–petrol, to give 1-(allyloxy)-2,2,6,6-tetramethylpiperidine **A5** (654 mg, 52%) as a colorless oil; ¹H NMR (400 MHz; CDCl₃) 5.91–5.81 (m, 1H), 5.27–5.21 (m, 1H), 5.10–5.06 (m, 1H), 4.27–4.25 (m, 2H), 1.56–1.28 (m, 6H), 1.15 (s, 6H), 1.10 (s, 6H); ¹³C NMR (100 MHz, CDCl₃) 134.2 (CH), 115.8 (CH₂), 78.4 (CH₂), 59.7 (C), 39.8 (CH₂), 33.0 (CH₃), 20.2 (CH₃), 17.3 (CH₂); HRMS *m/z* calculated for C₁₂H₂₄NO⁺ (M + H)⁺ 198.1852, found 198.1857 (+ 2.115 ppm error).

1-Benzyloxy-2,2,6,6-tetramethylpiperidine A6 A solution of 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO) (643 mg, 4.1 mmol) and naphthalene (24 mg, 0.19 mmol, 5 mol%) in THF (4 mL) was added to a Schlenk flask containing sodium (145 mg, 6.3 mmol, 1.5 eq.) in THF (4 mL). The reaction mixture was stirred at room temperature until the color changed from orange to dark green (1–2 h). The resulting solution was added to a solution of benzyl bromide (620 μL, 5.2 mmol, 1.3 eq.) and tetra-*n*-butylammonium iodide (300 mg, 0.93 mmol, 20 mol%) in THF (2 mL) and was stirred at room temperature for 18 hours. Diluted with saturated NH₄Cl solution (10 mL), extracted with

ethyl acetate (3 × 10 mL), dried over Na₂SO₄ and concentrated under reduced pressure. Purified by flash chromatography, eluting with 5% EtOAc–petrol, to give 1-benzyloxy-2,2,6,6-tetramethylpiperidine **A6** (940 mg, 92%) as a colorless oil; ¹H NMR (400 MHz; CDCl₃) 7.31–7.19 (m, 5H), 4.79 (s, 2H), 1.56–1.27 (m, 6H), 1.22 (s, 6H), 1.11 (s, 6H); ¹³C NMR (100 MHz, CDCl₃) 138.4 (C), 128.3 (CH), 127.5 (CH), 127.4 (CH), 78.9 (CH₂), 60.1 (C), 39.9 (CH₂), 33.2 (CH₃), 20.4 (CH₃), 17.3 (CH₂); HRMS *m/z* calculated for C₁₆H₂₆NO⁺ (M + H)⁺ 248.2009, found 248.2015 (+ 0.40 ppm error).

2,2,6,6-tetramethyl-1-(1-phenylethoxy)piperidine A7²⁶ A solution of 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO) (900 mg, 5.76 mmol) and CuBr (29 mg, 0.20 mmol, 4 mol%) in *t*-BuOH (4 mL) was stirred at room temperature. To this solution was added 2-phenylpropionaldehyde (1.6 g, 12 mmol) followed by slow addition of hydrogen peroxide (1.3 mL, 30% v/v, 17 mmol). The reaction mixture was heated to 50 °C for 2 hours. Saturated NH₄Cl solution (20 mL) was added, extracted with ethyl acetate (3 × 10 mL), dried over Na₂SO₄ and concentrated under reduced pressure. Purified by flash chromatography, eluting with 5% EtOAc–petrol, to give 2,2,6,6-tetramethyl-1-(1-phenylethoxy)piperidine **A7** (1.28 g, 85%) as a colorless solid; ¹H NMR (400 MHz; CDCl₃) 7.28–7.15 (m, 5H), 4.73 (q, 1H), 1.44–0.984 (m, 6H), 1.43 (d, 3H), 1.25 (s, 3H), 1.12 (s, 3H), 0.984 (s, 3H), 0.61 (s, 3H); ¹³C NMR (100 MHz; CDCl₃) 145.9 (C), 128.0 (CH), 126.8 (CH), 126.6 (CH), 83.2 (CH), 59.7 (C), 40.4 (CH₂), 34.5 (CH₃), 34.2 (CH₃), 23.6 (CH₃), 20.4 (CH₃), 17.3 (CH₂); HRMS *m/z* calculated for C₁₇H₂₈NO⁺ (M + H)⁺ 262.2165, found 262.2169 (– 0.80 ppm error).

2-((2,2,6,6-Tetramethylpiperidin-1-yl)oxy)propanenitrile A8 To a solution of 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO) (200 mg, 1.28 mmol), copper(I) bromide (220 mg, 1.54 mmol, 1.2 eq.) and *N,N,N',N'',N'''*-pentamethyldiethylenetriamine (320 μL, 1.53 mmol, 1.2 eq.) in toluene (4 mL) under N₂ was added 2-bromopropionitrile (110 μL, 1.53 mmol, 1.2 eq.) and the reaction mixture was heated to 50 °C for 30 minutes. Cooled and diluted with saturated NH₄Cl solution (20 mL), extracted with diethyl ether (3 × 20 mL), washed with brine (20 mL), dried over Na₂SO₄ and concentrated under reduced pressure. Purified by flash column chromatography, with petrol to 5% EtOAc–petrol as eluent, to give 2-((2,2,6,6-tetramethylpiperidin-1-yl)oxy)propanenitrile **A8** (250 mg, 93%) as a colorless oil; ¹H NMR (400 MHz; CDCl₃) 4.61 (q, *J* = 7.0 Hz, 1H), 1.53 (d, *J* = 7.0 Hz, 3H), 1.55–1.41 (m, 5H), 1.34–1.29 (m, 1H), 1.31 (s, 3H), 1.14 (s, 3H), 1.08 (s, 3H), 1.06 (s, 3H); ¹³C NMR (100 MHz, CDCl₃) 120.3 (C), 70.0 (CH), 60.8 (C), 59.8 (C), 39.9 (CH₂), 34.2 (CH₃), 33.6 (CH₃), 20.4 (CH₃), 20.2 (CH₃), 19.3 (CH₃), 17.0 (CH₂); HRMS *m/z* calculated for C₁₂H₂₃N₂O⁺ (M + H)⁺ 211.1805, found 211.1804 (– 0.662 ppm error).

Methyl 2-((2,2,6,6-tetramethylpiperidin-1-yl)oxy)acetate A9 A solution of 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO) (439 mg, 2.81 mmol) and naphthalene (20 mg, 0.16 mmol, 6 mol%) in THF (1 mL) was added to a Schlenk flask containing sodium (110 mg, 4.8 mmol, 1.7 eq.) in THF (1.5 mL). The reaction mixture was stirred at room temperature until the color changed from orange to dark green (1–2 h). This solution was added to a solution of methyl bromoacetate (300 μL, 3.2 mmol, 1.1 eq.) in THF (3 mL) and the reaction was stirred at room temperature for 18 hours. Diluted with saturated NH₄Cl solution (10 mL), and extracted with ethyl acetate (3 × 20 mL), washed with sat. NH₄Cl solution (3 × 20 mL) then brine (3 × 20

mL), dried over Na₂SO₄ and concentrated under reduced pressure. Purified by flash chromatography, eluting with 5% EtOAc–petrol to give methyl 2-((2,2,6,6-tetramethylpiperidin-1-yl)oxy)acetate **A9** (85 mg, 13%) as a colorless oil; ¹H NMR (400 MHz; CDCl₃) 4.44 (s, 2H), 3.73 (s, 3H), 1.45–1.26 (m, 6H), 1.14 (s, 12H); ¹³C NMR (100 MHz, CDCl₃) 170.1 (C), 75.4 (CH₂), 60.0 (C), 51.5 (CH₃), 39.7 (CH₂), 32.7 (CH₃), 20.1 (CH₃), 17.0 (CH₂); HRMS *m/z* calculated for C₁₂H₂₄NO₃⁺ (M + H)⁺ 230.1751, found 230.1751 (+ 0.347 ppm error).

Methyl 2-((2,2,6,6-tetramethylpiperidin-1-yl)oxy)propanoate A10 A solution of 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO) (435 mg, 2.78 mmol) and naphthalene (21 mg, 0.16 mmol, 6 mol%) in THF (1 mL) was added to a Schlenk flask containing sodium (98 mg, 4.3 mmol, 1.5 eq.) in THF (1.5 mL). The reaction mixture was stirred at room temperature until the color changed from orange to dark green (1–2 h). This solution was added to a solution of methyl 2-bromopropionate (350 μL, 3.1 mmol, 1.1 eq.) in THF (3 mL) and was stirred at room temperature for 18 hours. Diluted with saturated NH₄Cl solution (20 mL), and extracted with ethyl acetate (3 × 20 mL), washed with sat. NH₄Cl solution (3 × 20 mL) then brine (3 × 20 mL), dried over Na₂SO₄ and concentrated under reduced pressure. Purified by flash chromatography, eluting with 5% EtOAc–petrol to give methyl 2-((2,2,6,6-tetramethylpiperidin-1-yl)oxy)propanoate **A10** (488 mg, 72%) as a colorless oil; ¹H NMR (400 MHz; CDCl₃) 4.19 (q, *J* = 6.9 Hz, 1H), 3.57 (s, 3H), 1.43–1.15 (m, 6H), 1.26 (d, *J* = 7.0 Hz, 3H), 1.06 (s, 3H), 0.99 (s, 6H), 0.89 (s, 3H); ¹³C NMR (100 MHz, CDCl₃) 174.3 (C), 81.5 (CH), 59.9 (C), 59.4 (C), 51.2 (CH₃), 40.2 (CH₂), 40.0 (CH₂), 33.5 (CH₃), 32.8 (CH₃), 20.1 (CH₃), 19.9 (CH₃), 18.0 (CH₃), 17.0 (CH₂); HRMS *m/z* calculated for C₁₃H₂₆NO₃⁺ (M + H)⁺ 244.1907, found 244.1908 (+ 0.286 ppm error).

Methyl 2-methyl-2-((2,2,6,6-tetramethylpiperidin-1-yl)oxy)propanoate A11 A solution of 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO) (437 mg, 2.80 mmol) and naphthalene (22 mg, 0.17 mmol, 6 mol%) in THF (1 mL) was added to a Schlenk flask containing sodium (96 mg, 4.2 mmol, 1.5 eq.) in THF (1.5 mL). The reaction mixture was stirred at room temperature until the color changed from orange to dark green (1–2 h). The resulting solution was added to a solution of methyl 2-methyl-2-bromopropionate (400 μL, 3.09 mmol, 1.11 eq.) in THF (3 mL) and was stirred at room temperature for 18 hours. Diluted with saturated NH₄Cl solution (20 mL), and extracted with ethyl acetate (3 × 20 mL), washed with sat. NH₄Cl solution (3 × 20 mL) then brine (3 × 20 mL), dried over Na₂SO₄ and concentrated under reduced pressure. Purified by flash chromatography, eluting with 5% EtOAc–petrol to give 2-methyl-2-((2,2,6,6-tetramethylpiperidin-1-yl)oxy)propanoate **A11** (338 mg, 47%) as a colorless oil; ¹H NMR (400 MHz; CDCl₃) 3.72 (s, 3H), 1.63–1.26 (m, 6H), 1.47 (s, 6H), 1.14 (s, 6H), 0.99 (s, 6H); ¹³C NMR (100 MHz, CDCl₃) 176.6 (C), 81.2 (C), 59.6 (C), 51.8 (CH₃), 40.6 (CH₂), 33.4 (CH₃), 24.5 (CH₃), 20.5 (CH₃), 17.1 (CH₂); HRMS *m/z* calculated for C₁₄H₂₈NO₃⁺ (M + H)⁺ 258.2064, found 258.2064 (+ 0.077 ppm error).

2,2,6,6-tetramethyl-1-((tetrahydrofuran-2-yl)oxy)piperidine A12²⁷ To a suspension of 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO) (520 mg, 3.3 mmol) and copper(II) bromide (15 mg, 0.067 mmol, 2 mol%) in THF (2.7 mL, 33 mmol, 10 eq.) at 60 °C was added *tert*-butyl hydroperoxide solution (70% in H₂O, 1.3 mL, 10 mmol, 3 eq.) and the reaction mixture was stirred until the colour changed from orange to light green. The mixture was cooled, diluted with diethyl ether (20 mL), washed

with brine (2 × 20 mL), dried over Na₂SO₄ and concentrated under reduced pressure. Purified by column chromatography, with 10% EtOAc–petrol as eluent, to give 2,2,6,6-tetramethyl-1-((tetrahydrofuran-2-yl)oxy)piperidine **A12** (610 mg, 81%) as a colourless oil; ¹H NMR (400 MHz; CDCl₃) 5.27–5.25 (m, 1H), 3.80–3.69 (m, 2H), 1.93–1.79 (m, 3H), 1.74–1.63 (m, 1H), 1.50–1.34 (m, 5H), 1.28–1.19 (m, 1H), 1.13 (s, 3H), 1.02 (s, 3H), 0.98 (s, 3H), 0.95 (s, 3H); ¹³C NMR (100 MHz, CDCl₃) 109.6 (CH), 66.5 (CH₂), 60.0 (C), 58.5 (C), 40.1 (CH₂), 39.7 (CH₂), 33.9 (CH₃), 33.3 (CH₃), 31.2 (CH₂), 23.9 (CH₂), 20.4 (CH₃), 20.0 (CH₃), 17.2 (CH₂); HRMS *m/z* calculated for C₁₃H₂₆NO₂⁺ (M + H)⁺ 228.1958, found 228.1960 (+ 0.896 ppm error).

2,2,6,6-tetramethylpiperidin-1-yl acetate A13 A solution of TEMPO (740 mg, 4.7 mmol) and naphthalene (32 mg, 0.25 mmol, 5 mol%) in THF (4 mL) was added to a Schlenk flask containing sodium (151 mg, 6.6 mmol, 1.5 eq.) in THF (4 mL). The reaction mixture was stirred at room temperature until the color changed from orange to dark green (1–2 h). The resulting solution was added to a solution of acetyl chloride (430 μL, 6.0 mmol, 1.3 eq.) in THF (2 mL) at 0 °C, then stirred at room temperature for 18 hours. Diluted with saturated NH₄Cl solution (10 mL), extracted with ethyl acetate (3 × 10 mL), dried over Na₂SO₄ and concentrated under reduced pressure. Purified by flash chromatography, eluting with 10% EtOAc–petrol, to give 2,2,6,6-tetramethylpiperidin-1-yl acetate **A13** (475 mg, 50%) as a white solid; ¹H NMR (400 MHz; CDCl₃) 2.05 (s, 3H), 1.70–1.36 (m, 6H), 1.12 (s, 6H), 1.03 (s, 6H); ¹³C NMR (100 MHz, CDCl₃) 170.9 (C), 60.0 (C), 39.0 (CH₂), 32.0 (CH₃), 20.5 (CH₃), 19.2 (CH₃), 17.0 (CH₂); HRMS *m/z* calculated for C₁₁H₂₂NO₂⁺ (M + H)⁺ 200.1645, found 200.1647 (+ 0.772 ppm error).

2,2,6,6-tetramethyl-1-phenoxypiperidine A14²⁸ To a solution of 2,2,6,6-tetramethyl-1-piperidinyloxy (TEMPO) (1.0 g, 6.4 mmol), copper(II) fluoride (6.5 mg, 0.064 mmol, 1 mol%) and *tert*-butyl nitrite (1.58 mL, 13.3 mmol, 2 eq.) in MeCN (25 mL) at 65 °C was added aniline (1.22 mL, 13.4 mmol, 2.1 eq.) and the reaction mixture was stirred for 10 minutes, then allowed to cool. Diluted with brine (20 mL), extracted with diethyl ether (3 × 20 mL), dried over Na₂SO₄ and concentrated under reduced pressure. Purified by flash column chromatography, with 5% EtOAc–petrol as eluent, to give 2,2,6,6-tetramethyl-1-phenoxypiperidine **A14** (584 mg, 39%) as a colourless oil; ¹H NMR (400 MHz; CDCl₃) 7.15–7.09 (m, 4H), 6.78–6.74 (m, 1H), 1.61–1.45 (m, 5H), 1.38–1.30 (m, 1H), 1.16 (s, 6H), 0.94 (s, 6H); ¹³C NMR (100 MHz, CDCl₃) 163.7 (C), 128.8 (CH), 120.0 (CH), 114.0 (CH), 60.4 (C), 39.9 (CH₂), 32.8 (CH₃), 20.6 (CH₃), 17.2 (CH₂); HRMS *m/z* calculated for C₁₅H₂₄NO⁺ (M + H)⁺ 234.1852, found 234.1862 (+ 1.70 ppm error).

RESULTS AND DISCUSSION

Cyclic Voltammetry. The alkoxyamines shown in Figure 2 were synthesized and then studied by cyclic voltammetry. Representative cyclic voltammograms under consistent conditions are shown in Figure 3. For clarity, Figure 3 shows the second scan only for each compound (i.e. immediately following an initial scan under the same conditions). Where the second scan shows irreversible alkoxyamine oxidation and reversible nitroxide oxidation, the first scan resembles that measured previously for TEMPO-CH(CH₃)Ph (see Figure 1).⁸ That is, the first scan shows irreversible alkoxyamine oxidation, no nitroxide oxida-

tion wave but an oxoammonium reduction wave, thus consistent with oxidative cleavage of the alkoxyamine via an EC_{irrev}E mechanism to a nitroxide and carbocation or an EC_{irrev}E (or EC_{irrev}E) mechanism to an oxoammonium and carbon-centered radical. Complete CV results are provided in the Supporting Information.

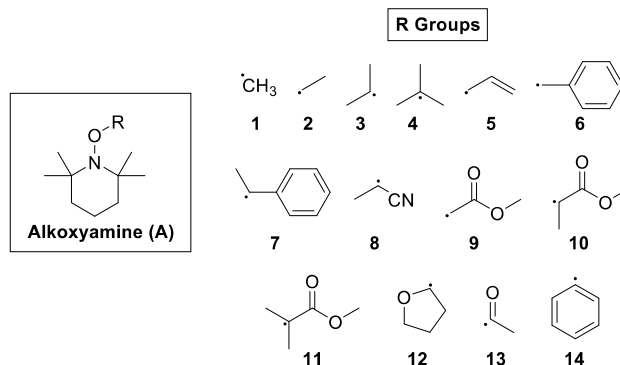


Figure 2. Test set of alkoxyamines (**A1–A14**) studied in this work.

Investigation of TEMPO-alkyls (**A1–A4**, R = Me, Et, *i*-Pr and *t*-Bu) revealed alkoxyamine oxidation occurs at similar potentials to form the corresponding radical-cation intermediate (TEMPO-R → [TEMPO-R]^{•+}). However, the cleavage patterns differed quite substantially. TEMPO-Me (**A1**) and TEMPO-Et (**A2**) showed reversible alkoxyamine oxidation and hence no fragmentation (Figure 3A). In contrast, oxidation of TEMPO-*t*-Bu (**A4**) was completely irreversible and subsequently cleaved to produce a second peak corresponding to TEMPO[•]/TEMPO⁺ (Figure 3B). TEMPO-*i*-Pr (**A3**) appeared to have partial alkoxyamine reversibility whilst continuing to cleave to form TEMPO[•]/TEMPO⁺ (Figure 3B). These results suggest that alkoxyamine cleavage is favored by the presence of a secondary or tertiary carbon in the α-position to the central N–O bond. This is intuitive in that additional α-methyl groups can hyperconjugatively stabilize the leaving carbocation (or radical) with stabilities in the order Me < Et < *i*-Pr < *t*-Bu. In a similar vein, the allylic and benzylic derivatives TEMPO-allyl (**A5**), TEMPO-Bn (**A6**) and TEMPO-CH(CH₃)Ph (**A7**) oxidize irreversibly and all show alkoxyamine fragmentation to produce TEMPO[•]/TEMPO⁺ and the carbocation or carbon-centered radical (Figure 3C). This is consistent with the ability of these groups to afford resonance stabilization of the R-fragment.

Electron withdrawing groups destabilize carbocations and thus reversible oxidation with no cleavage might be expected with these derivatives. However, the nitrile compound (**A8**) shows complete irreversibility, albeit at a higher oxidation potential (Figure 3D). For the methyl ester analogues (**A9–A11**) however, a similar trend was observed to the unsubstituted alkyl groups: from primary to tertiary substitution at the α position (to the N–O bond) of the ester, behavior changes from partial reversibility to irreversibility for alkoxyamine oxidation (Figure 3E). As these electron withdrawing groups are capable of stabilizing radicals but not carbocations, these results suggest that the oxidative cleavage observed may be the result of cleavage to a radical instead, a result supported by theoretical calculations (*vide infra*).

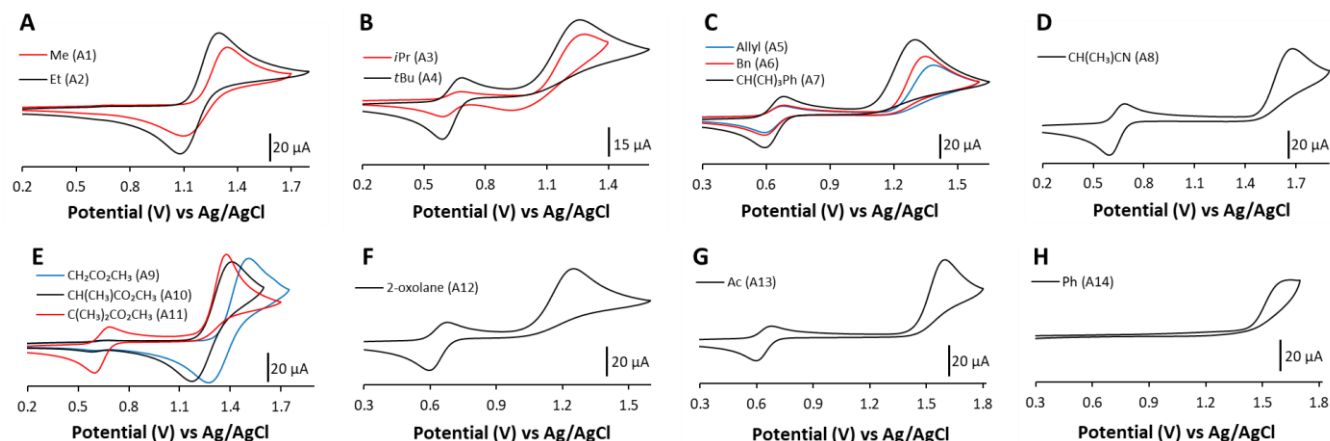


Figure 3. The second scan of cyclic voltammograms of various alkoxyamines with R-groups containing: **A** and **B**, alkyl derivatives, **C**, allylic and benzylic derivatives, **D**, nitrile substitution, **E**, methyl ester substitution, **F**, 2-oxolane, **G**, acetyl, **H**, phenyl. Alkoxyamine (2 mM) in Bu₄NClO₄/MeCN (0.1 M), sweep rate 250 mV s⁻¹. Experimental conditions and voltammograms over a range of sweep rates can be found in the experimental section and in the Supporting Information.

Following this, the effect of adjacent heteroatoms was investigated. Unsurprisingly, TEMPO-2-oxolane (**A12**) was irreversibly oxidized and cleaved consistent with the expected stabilization of the cation by lone pair donation from the adjacent oxygen (Figure 3F). Similarly, TEMPO-Ac (**A13**) exhibited oxidative cleavage, promoted by electron delocalization in the fragment acylium species (Figure 3G).

Finally, we aimed to investigate the oxidation and cleavage patterns of a TEMPO-based alkoxyamine that would produce an unstable fragment irrespective of the cleavage pathway (an unstable radical or an unstable cation). Surprisingly, TEMPO-Ph (**A14**) also underwent irreversible oxidation of the parent alkoxyamine but, unlike the other systems studied, showed no production of TEMPO[•]/TEMPO⁺ as a fragment (Figure 3H). We propose that cleavage does in fact occur, albeit by a different mechanism (*vide infra*).

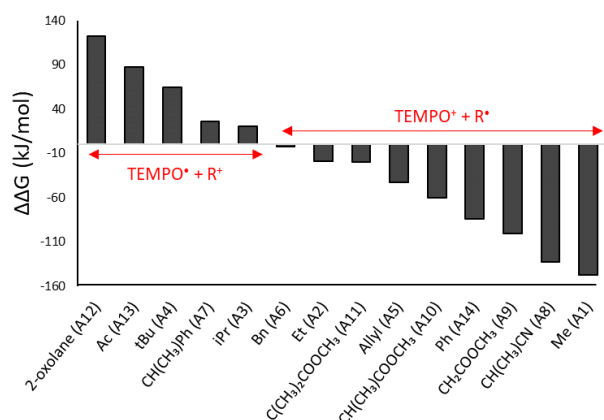


Figure 4. Difference in Gibbs free energies for oxidative cleavage of TEMPO-R to TEMPO[•] + R[•] versus TEMPO[•] + R[•] in acetonitrile at 298.15 K. A positive number indicates cleavage to the carbocation is preferred while a negative number indicates that cleavage to a carbon-centered radical is preferred.

Cleavage Pathways. As in our previous work,⁸ accurate computational chemistry was used to discriminate between the

carbocation and carbon-centered radical cleavage pathways of Scheme 2. Figure 4 shows the difference in Gibbs free energies for each pathway, defined such that a positive number indicates that the carbocation is favored. Unsurprisingly, R-groups with electron-donating substituents favor the carbocation pathway and those with electron withdrawing groups favor the radical pathway. Species that are not stabilized as radicals or cations (Me, Ph) also favor the radical pathway in principle, although in the case of TEMPO-Me cleavage does not occur at all.

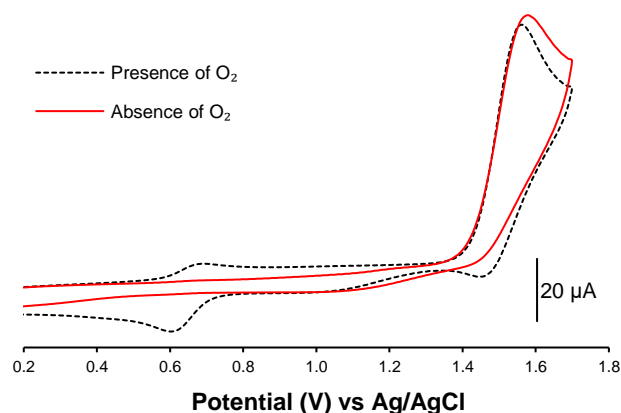
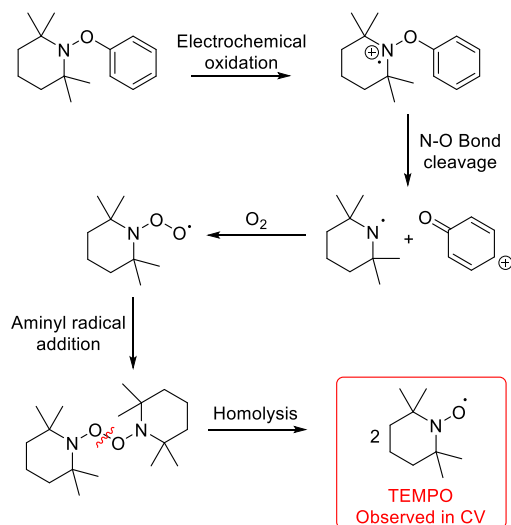


Figure 5: Cyclic voltammogram indicating N–O cleavage of TEMPO-Ph (**A14**). TEMPO is observed only when conducted open to the air. TEMPO-Ph (1.6×10^{-3} M) in Bu₄NClO₄/MeCN (0.1 M), sweep rate 1000 mV s⁻¹.

As noted above, TEMPO-Ph (**A14**) shows unusual behavior compared with the rest of the test set. It clearly undergoes irreversible oxidation and hence some irreversible chemical reaction but the TEMPO[•]/TEMPO⁺ redox couple is not formed. While most neutral alkoxyamines undergo homolytic cleavage at the C–O bond, when the leaving radical is particularly unstable N–O cleavage can occur instead.²⁹ We therefore suspected that N–O cleavage of the oxidized alkoxyamine may explain the unusual behavior of the TEMPO-Ph species. It is known from work focused on the cleavage of neutral alkoxyamines in the context of hindered amine light stabilizer chemistry that oxygen

can trap aminyl radicals to later form air-stable nitroxides.³⁰ To probe this reactivity experimentally, oxygen was reintroduced into the system when conducting the CV (Figure 5). As a result, nitroxide formation (TEMPO) was observed in the presence of oxygen thus confirming the production of aminyl radicals and consequently the N–O cleavage mode of fragmentation (Scheme 3).



Scheme 3: Proposed mechanism of TEMPO-Ph (**A14**) cleavage in the presence of O₂.

Oxidation Potentials. Finally, the cyclic voltammetry data were fitted to obtain accurate estimates of the oxidation potentials and to help confirm the cleavage mechanism (Appendix

S3). For R = Me and Et, the data were fitted with a simple reversible oxidation wave. For the others, the mechanism was more complicated. Three different processes were considered:

1. EC_{irrev}E where TEMPO-R⁺⁺ undergoes fragmentation with a rate coefficient (k_{frag}) to either (a) R⁺ and TEMPO[•] with TEMPO[•] then oxidizing, or (b) R[•] and TEMPO⁺ with R[•] then oxidizing. The R⁺ in either case is assumed to be electrochemically inactive due to association with ClO₄[−].
2. EC_{irrev}E where TEMPO-R⁺⁺ undergoes fragmentation with a rate coefficient (k_{frag}) to TEMPO⁺ but R[•] does not undergo further oxidation as its oxidation potential is greater than that of TEMPO-R. This is often described as “parallel ECE” as the second electrochemical step (oxoammonium reduction) occurs in the opposite direction to the first, but still within the potential window of the experiment.³¹
3. EC_{irrev} where the initial oxidation leads to electrochemically inactive products. This would be consistent with the N–O fragmentation process in Scheme 3, but also with other chemical decomposition pathways of the alkoxyamine such as deprotonation and subsequent β-scission.

While pathways 1a and 1b cannot be discriminated based on fitting alone, the computational results of Figure 4, coupled with the computational oxidation potentials of R[•] can be used to establish whether pathway 1a or 1b is likely. Table 1 shows the fitted and quantum-chemical oxidation potentials for test set, along with the quantum-chemical oxidation potentials of R[•] and pathway that provides the best fit to the data, and hence the most likely cleavage mechanism.

Table 1. Experimental^a and theoretical^b oxidation potentials (V vs Ag/AgCl in acetonitrile) for TEMPO-R and R[•], the best-fit mechanism for the CV data, the best-fit fragmentation rate coefficient (k_{frag} , where relevant) and the most likely cleavage mechanism.

TEMPO-R R-group	Theory		Experiment			Cleavage Mechanism ^f
	R [•]	TEMPO-R	TEMPO-R ^c	Fitted	k_{frag} (s ^{−1})	
A1 Me	2.13	1.31	1.22 ± 0.00	E	–	none
A2 Et	0.82	1.31	1.18 ± 0.00	E	–	none
A3 <i>i</i> -Pr	0.43	1.17	1.10 ± 0.01	EC _{irrev} E (1)	0.31	TEMPO [•] / R ⁺
A4 <i>t</i> -Bu	−0.05	1.25	1.25 ± 0.03	EC _{irrev} E (1)	>10 ⁶	TEMPO [•] / R ⁺
A5 allyl	0.79	1.35	1.37 ± 0.04	EC _{irrev} E (1) ^{d,e}	1400	TEMPO ⁺ / R [•]
A6 Bn	0.65	1.35	1.32 ± 0.00	EC _{irrev} E (1) ^d	1800	both possible
A7 CH(CH ₃)Ph	0.36	1.26	1.26 ± 0.02	EC _{irrev} E (1)	12500	TEMPO [•] / R ⁺
A8 CH(CH ₃)CN	2.00	1.62	1.63 ± 0.01	EC _{irrev} E (2)	1015	TEMPO [•] / R [•]
A9 CH ₂ C(O)OCH ₃	1.64	1.46	1.38 ± 0.00	EC _{irrev} E (2)	0.01	TEMPO ⁺ / R [•]
A10 CH(CH ₃)C(O)OCH ₃	1.50	1.33	1.30 ± 0.00	EC _{irrev} E (2)	0.01	TEMPO ⁺ / R [•]
A11 C(CH ₃) ₂ C(O)OCH ₃	1.24	1.41	1.37 ± 0.01	EC _{irrev} E (1)	>10 ⁶	TEMPO ⁺ / R [•]
A12 2-oxolane	−0.64	1.29	1.22 ± 0.00	EC _{irrev} E (1)	> 10 ⁶	TEMPO [•] / R ⁺
A13 Ac	−0.28	1.61	1.60 ± 0.01	EC _{irrev} E (1)	>10 ⁶	TEMPO [•] / R ⁺
A14 Ph	1.50	1.62	1.57 ± 0.08	EC _{irrev} (3)	390	N–O

^aElectrochemical formal potential of the alkoxyamine/alkoxyamine⁺⁺ redox couple, as estimated from digital simulations of experimental cyclic voltammetry data. The experimental oxidation potential of TEMPO itself was 0.64 V. ^bCalculated using a 2 layer ONIOM procedure with G3(MP2)CC(+)//M06-2X/6-31+G(d,p) for the core and G3MP2CC//M06-2X/6-31+G(d,p) for the outer layer. Solvation was calculated with COSMO-RS. ^c Upper and lower limits of the 95.4% confidence level are in Fig. S1 to S14 ^dA contribution by an EC_{irrev} process of comparable magnitude to the respective mesolysis step improved the fits for these reactions. ^e For this one, the best fit was obtained when the oxidation of R[•] was assumed to be reversible. ^f Based on the fitting and computational results..

From Table 1, the first observation is that for the reactions fitted by pathway 2 (i.e., fragmentation to an R[•] that does not oxidize in the experimental window), the theoretical R[•] oxidation potential is indeed larger than that of R-TEMPO and thus inaccessible, consistent with the mechanism fitted. For the reactions fitted by pathway 1, most are expected to produce carbocations (see Figure 4), consistent with pathway 1a. The main exceptions are R = C(CH₃)₂C(O)OCH₃ and allyl, though R = Bn shows no distinct preference either way. In these three cases the R[•] oxidation potential is indeed less than that of R-TEMPO, as required for pathway 1b. In other words, the cleavage mechanisms predicted by theory are fully consistent with the mechanisms fitted to the cyclic voltammograms.

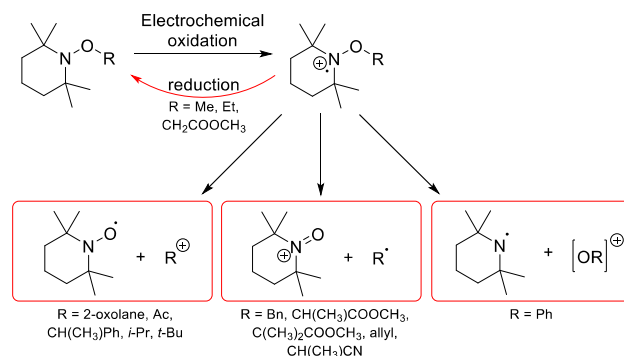
The second observation is that, where fragmentation of the oxidized alkoxyamine occurs, the rate coefficients span a considerable magnitude. Within the R = Me, Et, *i*-Pr, *t*-Bu series, the rate coefficients range from zero (Me, Et) to 10⁻¹ s⁻¹ (*i*-Pr) to > 10⁶ s⁻¹ (*t*-Bu). The other fragmentations to carbocations (R = CH(CH₃)Ph, 2-oxolane, Ac) are also fast > 10⁴ s⁻¹, and thus these alkoxyamines show promise as sources of carbocations for synthesis. In contrast, the fragmentations to radicals are generally very slow (10⁻² – 10² s⁻¹) and accompanied by side reactions, with only R = C(CH₃)₂C(O)OCH₃ and allyl showing potential as a possible clean source of radicals. For R = C(CH₃)₂C(O)OCH₃ fragmentation is fast (*k_f* > 10⁶) oxidation to R⁺ does not occur in the experimental window. For R = allyl oxidation does occur but importantly the oxidation of R[•] to R⁺ is reversible, presumably due to the resonance stabilization of the allyl carbocation.

Finally, we note that the experimental oxidation potentials of the alkoxyamines are relatively high, falling into the range of 1.1–1.6 V versus Ag/AgCl. As a result, the reversible systems (e.g. R = Me, Et) would be promising in battery applications. The oxidation potentials show good correlation with the Hammett σ_R parameters of both the R-group (*R*² = 0.76) and OR-group (*R*² = 0.88) of the alkoxyamine (see Supporting Information Figure S18). The better correlation with the OR Hammett parameters is due to the fact that the radical-cation character of the oxidized species is on the N rather than O atom, and thus it is better to view this substituent effect as OR stabilization of N⁺ rather than R stabilization of the oxidized N–O⁺. Comparing theory and experiment, we note that the theoretical oxidation potential of TEMPO is 0.62 V, in good agreement with experiment (0.64 V). Likewise, where experimental data was available from the literature,³² the theoretical R[•] oxidation potentials (R[•] = *t*-Bu, -0.05 V; Bn, 0.65 V; CH(CH₃)Ph, 0.36 V) are in good agreement (MAD 0.06 V) with experiment (0.045 V, 0.685 V and 0.325 V, respectively). For the alkoxyamines the deviations between theory and experiment (Table 1) are of a similar magnitude (MAD 0.04 V), thus mutually confirming the accuracy of the calculations and mechanisms fitted.

CONCLUSION

Electrochemical methods are undergoing a renaissance in organic synthesis due to their atom efficiency, low cost and functional group tolerance. Here we have shown that TEMPO-based alkoxyamines can provide a mild electrochemical source of carbon-centered radicals or carbocations depending on the stability of the respective leaving group (Scheme 4). Although the nature of the species generated (R⁺ versus R[•]) is limited by the relative stabilities of the carbocation and radical, both are useful in synthesis and can be difficult to generate by other means without

harsh conditions and/or toxic reagents. Moreover, because alkoxyamines can be tethered to surfaces, their potential applications extend beyond synthesis to surface modification.⁹ Investigation of these synthetic applications is currently underway.



Scheme 4. Cleavage pathways of the test set.

ASSOCIATED CONTENT

Supporting Information

¹H and ¹³C NMR spectra, complete cyclic voltammetry data, kinetic modelling and computational details. The Supporting Information is available free of charge on the ACS Publications website.

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Author Contributions

The manuscript was written through contributions of all authors. / All authors have given approval to the final version of the manuscript.

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